

MAHESH PUBLIC SCHOOL, JODHPUR

REVISION NOTES - 1

CLASS IX

Subject : SCIENCE

Chapter : Matter In Our Surroundings

Change of State of Matter:

Physical states of matter can be interconverted into each other by following two ways:

1. By changing the temperature

2. By changing the pressure

1. Effect of Change of Temperature:

Solid to liquid:

- On increasing the temperature of solids, the kinetic energy of the particles increases which overcomes the forces of attraction between the particles thereby solid gets converted to a liquid.
- **Melting:** Change of solid state of a substance into liquid is called melting.
- **Melting point:** The temperature at which a solid melts to become a liquid at the atmospheric pressure is called its melting point.
- Melting point of ice is 0°C.

(a) Liquid to gas:

- On heating a liquid like water, the kinetic energy of its particles increases as high as in a gas, thus causing the liquid to change to a gas.
- **Boiling:** The change of a liquid substance into gas on heating is called boiling.
- **Boiling point:** The temperature at which a liquid boils and changes rapidly into a gas at the atmospheric pressure is called its boiling point.
- Boiling point of water is 100°C.

(b) Gas to liquid:

- On cooling a gas like steam (or water vapour), the kinetic energy of its particles is lowered down, causing them to move slowly and bringing them closer, forming a liquid.
- **Condensation:** The process, in which a gas, on cooling, turns into a liquid at a specific temperature is called condensation or liquefaction.

(c) Liquid to solid:

- When a liquid is cooled down by lowering its temperature, its particles lose the kinetic energy and come to a stationary position, causing the liquid to turn to solid.
- **Freezing:** The change of a liquid substance into solid by lowering its temperature is called freezing.
- **Freezing point:** The temperature at which the state of a substance changes from a liquid to a solid is called the freezing point of that substance.

Fusion: The process of melting, that is, change of solid state into liquid state is also known as fusion.

Latent heat: The heat energy that is required to change the state of a substance without causing any rise in the temperature of the substance is called latent heat. Since, the heat energy is hidden in the bulk of the matter, it is called latent heat.

- **Latent heat of fusion:** The heat energy required to convert 1 kilogram of a solid into liquid at atmospheric pressure, at its melting point, is known as the latent heat of fusion
- **Latent heat of vaporisation:** The heat energy required to convert 1 kilogram of liquid into gas, at atmospheric pressure, at its boiling point, is known as the latent heat of vaporisation

Note: *Water vapour at 373 K have more energy than water at the same temperature because particles in steam have absorbed extra energy in the form of latent heat of vaporisation.*

Sublimation: The change of state of a substance directly from a solid to gas or gas to solid, without changing into the liquid state, is called sublimation.

2. Effect of change of pressure

- **Gas to liquid:** Gases can be liquefied by applying pressure and reducing the temperature. When a high pressure is applied to a gas, it gets compressed and if the temperature is lowered, the gas is liquefied.
- Solid CO₂ gets converted directly to gaseous state on decrease of pressure to 1 atmosphere without coming into liquid state. This is the reason that solid carbon dioxide is also known as dry ice.

Evaporation:

The process of conversion of a substance from the liquid state to the gaseous state at any temperature below its boiling point is called evaporation or vaporisation.

Factors affecting the rate of evaporation:

- Surface area: The rate of evaporation increases on increasing the surface area of the liquid.
- Temperature: The rate of evaporation increases with an increase in temperature.
- Humidity: Decrease in the humidity increases the rate of evaporation.
- Wind speed: An increase in the wind speed increases the rate of evaporation.

Evaporation causes cooling:

During the process of evaporation, the particles of liquid absorb energy or latent heat of vaporisation from the surrounding to get converted to gaseous state. This absorption of energy from the surroundings make the surroundings cold.

For example: The perspiration or sweating in our body keep the body temperature constant by taking away the extra heat from body as the latent heat of vaporisation.